

For the decomposition reaction, $2\text{NO}_2 (\text{g}) \rightarrow 2\text{NO} (\text{g}) + \text{O}_2 (\text{g})$ @ 300°C, using the data from the table below make five computer generated plots.

- $[\text{NO}_2]$ vs time
- $[\text{NO}]$ vs time
- $[\text{O}_2]$ vs time
- $\ln[\text{NO}_2]$ vs time
- $1/[\text{NO}_2]$ vs time

Time (± 1 s)	NO_2	NO	O_2
0	0.0100	0	0
50	0.0079	0.0021	0.0011
100	0.0065	0.0035	0.0018
150	0.0055	0.0045	0.0023
200	0.0048	0.0052	0.0026
250	0.0043	0.0057	0.0029
300	0.0038	0.0062	0.0031
350	0.0034	0.0066	0.0033
400	0.0031	0.0069	0.0035