Quiz 9

Which of the following solvents could be miscible with H₂O?

\[ \text{CH}_2\text{Cl}_2, \text{CCl}_4, \text{CH}_3\text{CH}_2\text{OH}, \text{propane (C}_3\text{H}_8) \]

- CH₂Cl₂ - polar molecule
- CCl₄ - nonpolar
- CH₃CH₂OH - polar & H-bonding
- C₃H₈ - nonpolar

\( \text{CH}_3\text{CH}_2\text{OH} \) is definitely soluble in H₂O (OH group)

CH₂Cl₂ might be soluble in H₂O because it is polar.

Note: in "real life", CH₂Cl₂ is not soluble in H₂O because it isn't strongly polar & has no H-bonding.