Chapter 6 (for Exam 3)

- sign conventions for ΔE, ΔH, q, w
- $\Delta E = q + w$
- state function
- calorimetry: bomb and "coffee cup"
  - (ΔE) ΔH
  - $\Delta V = \emptyset$ $(\Delta P = \emptyset)$

heat involved = heat gained by + heat gained by water or solution + by calorimeter

- Hess' Law : 2 ways to calculate
  - (1) manipulating 2 or 3 equations
  - (2) using ΔH_f°'s

  $$\Delta H_{rxn} = \sum n \Delta H_f^\circ \text{ products} - \sum n \Delta H_f^\circ \text{ reactants}$$

- meaning of ΔH_f°
- using heat involved & amt of material involved