Examples of Inorganic Nomenclature and Charges of Ions:

CATIONS		ANIONS	
Formula	**Name**	**Formula**	**Name**
Li^{+}	lithium ion	N^{3-}	nitride ion
Na^{+}	sodium ion	P^{3-}	phosphide ion
K^{+}	potassium ion	O^{2-}	oxide ion
Rb^{+}	rubidium ion	S^{2-}	sulfide ion
Cs^{+}	cesium ion	Se^{2-}	selenide ion
Mg^{2+}	magnesium ion	F^{-}	fluoride ion
Ca^{2+}	calcium ion	Cl^{-}	chloride ion
Sr^{2+}	strontium ion	Br^{-}	bromide ion
Ba^{2+}	barium ion	I^{-}	iodide ion
Cr^{2+}	chromium(II) ion	CN^{-}	cyanide ion
Cr^{3+}	chromium(III) ion	OH^{-}	hydroxide ion
Mn^{2+}	manganese(II) ion	CO_3^{2-}	carbonate ion
Fe^{2+}	iron(II) ion	HCO_3^{-}	hydrogen carbonate ion (bicarbonate ion)
Fe^{3+}	iron(III) ion	NO_3^{-}	nitrate ion
Co^{2+}	cobalt(II) ion	NO_2^{-}	nitrite ion
Co^{3+}	cobalt(III) ion	PO_4^{3-}	phosphate ion
Ni^{2+}	nickel(II) ion	HPO_4^{2-}	hydrogen phosphate ion
Cu^{+}	copper(I) ion	$H_2PO_4^{-}$	dihydrogen phosphate ion
Cu^{2+}	copper(II) ion	PO_3^{3-}	phosphite ion
Sn^{2+}	tin(II) ion	AsO_4^{3-}	arsenate ion
Sn^{4+}	tin(IV) ion	SO_4^{2-}	sulfate ion
Zn^{2+}	zinc ion*	HSO_4^{-}	hydrogen sulfate ion
Al^{3+}	aluminum ion*	SO_3^{2-}	sulfite ion
Ag^{+}	silver ion*	ClO_4^{-}	perchlorate ion
Cd^{2+}	cadmium ion*	ClO_3^{-}	chlorate ion
NH_4^{+}	ammonium ion	ClO_2^{-}	chlorite ion
**$CH_3NH_3^{+}$	methylammonium ion**	ClO^{-}	hypochlorite ion
**$(CH_3)_2NH_2^{+}$	dimethylammonium ion**	BrO_4^{-}	perbromate ion
**$(CH_3)_3NH^{+}$	trimethylammonium ion**	BrO_3^{-}	bromate ion
		BrO_2^{-}	bromite ion
		BrO^{-}	hypobromite ion
		IO_4^{-}	periodate ion
		IO_3^{-}	iodate ion
		IO_2^{-}	iodite ion
		IO^{-}	hypoiodite ion
		CH_3COO^{-}	acetate ion
		$CHOO^{-}$	formate ion
		MnO_4^{-}	permanganate ion
		CrO_4^{2-}	chromate ion
		$Cr_2O_7^{2-}$	dichromate ion

* no Roman numerals a necessary since the ion only exists in one oxidation state

** for Chem 102

Required Weak Acids and Weak Base Nomenclature for Chem 102:

(See notes for required acids and bases for Chem 101)

WEAK ACIDS		**WEAK BASES**	
Formula	**Name**	**Formula**	**Name**
CH_3COOH	acetic acid	NH_3	ammonia
CHOOH	formic acid	CH_3NH_2	methylamine
HCN(aq)	hydrocyanic acid	$(CH_3)_2NH$	dimethylamine
HF(aq)	hydrofluoric acid	$(CH_3)_3N$	trimethylamine
HClO	hypochlorous acid		
HBrO	hypobromous acid		
HNO_2	nitrous acid		

For interest: