

Chapter 18 : Slightly Soluble Compounds (for Exam 3)

- solubility product expression and K_{sp} (mol/L)
- calculations with molar solubility, S , and solubility (g/100mL, pH, cmcⁿ)
- common ion effect - Le Chatelier's Principle theory + calculations

For Final

- From K_{sp} vs Q_{sp} , will a precipitate form when 2 solutions are mixed.

$Q_{sp} < K_{sp}$	no
$Q_{sp} > K_{sp}$	yes
- What concentration of ion is required to initiate precipitation of a solid - solve K_{sp} equation